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REVIEW PROPER 4. UNSATURATED HYDROCARBONS (ALKANES AND ALKYNES)

PART 1: INTRODUCTION

Unsaturated hydrocarbons are not completely saturated with hydrogens due to presence of pi bonds (double or triple). Those with double bonds are called alkenes, and those with triple bonds are called alkynes.

Alkenes—one or more carbon-carbon double bonds.

A linear alkene has the general formula of C_nH_{2n} .

- **Vinyl carbons** - carbon atoms directly double bonded to each other. Hydrogens attached to it are called vinylic or alpha hydrogens.
- **Alkyl carbons** - carbons directly bonded to vinyl carbons.

The alkene carbons bearing the double bond (except cumulated double) have trigonal planar geometry by being sp², and cumulated double have linear geometry by being sp hybridization of carbon.

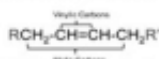


Figure 4.1. The vinyl and alkyl carbons of an alkene.

Alkynes—one or more carbon-carbon triple bonds.

A linear alkyne has the general formula of C_nH_{2n-2} .

All alkyne carbons bearing the triple bond have linear geometry. A linear shape has no edge, so a linear formula has no bonds.



Figures 4.2 and 4.3. The skeletal formulas of the bottom are initially awkward to see, but these are the correct ways of drawing an sp hybridized carbon.

Enyne—a hydrocarbon with at least one double bond and one triple bond.

PART 2: PROPERTIES OF ALKENES

- are non-polar, and liquid at room temperature.
- have densities of less than 1 g/cm³.
- are more reactive than alkanes and alkynes because of their diffuse electron clouds. Alkyne's reactivity are weakened due to overcrowding of electrons in the triple bond.
- **Trans isomers** are more stable than cis isomers in simple alkenes (due to less steric effect)
 - $n \cdot 2^n$ for male to get how many isomers where "n" is the number of double bonds.
- The more substituents present around the vinyl carbon, the more stable the molecule (alkyl).
- **Internal alkenes/alkynes** (those with double/triple bonds inside the parent chain) are more stable than terminal alkenes (those with double/triple bonds at the end/terminus of the parent chain). (Overall structural effects to answer why).

INDEX OF HYDROGEN DEFICIENCY (IHD)

- gives the number of H₂ molecules needed to fully saturate a hydrocarbon; in the laboratory, IHD of alkenes/alkynes is called the degree of unsaturation.
- IHD can help in predicting the structure of a hydrocarbon that may possess a ring or pi bonds (or both).
- Notice that for every pi bond or ring closure, TWO hydrogens are removed from the compound. This is why each IHD value corresponds to TWO hydrogen atoms.

Steps for calculating IHD:

1. Count the number of carbon and hydrogen atoms of given, then write its molecular formula (C_nH_m).
2. Using the given's carbons (n), compute for its linear alkane formula using C_nH_{2n+2}.
3. Subtract the hydrogens from given to hydrogens of the linear alkane.
4. Divide the difference by 2 since one IHD value corresponds to two hydrogens. The quotient will give you the IHD.



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